

# task\_4pla882qt21jfyc9\_with\_calculation

## Student Group

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[charges, Electroplating, chapter1 2](#)**Exercise E18 Charges in Electroplating**

To get a different metal coating onto a surface, often [Electroplating](#) is used. In this process, the surface is located in a liquid, which contains metal ions of the coating.

In the following, a copper coating (e.g. for corrosion resistance) shall be looked on. The charge of one copper ion is around  $1.6022 \cdot 10^{-19} \text{ C}$ , what is the charge on the surface if there are  $8 \cdot 10^{22}$  ions added?

$$\begin{aligned} & 12'818 \text{ C} \end{aligned}$$

Solution

$$\begin{aligned} & 8 \cdot 10^{22} \cdot 1.6022 \cdot 10^{-19} \text{ C} = 12'817.6 \text{ C} \end{aligned}$$

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